	Paper II-	- Answer Sheet	
1. (A) i).	2.000		
	2 000	2 000	
	Producers Primary Consumers	Primary Productions	cers
	For one of the above graphs		- 02
ii)	x = 200 J(Without the unit 01 mark)		- 02
-	Respiration / energy is spent on biological		- 01
=	The balanced relationship between organis a – burning of fossil fuel / respiration b- decomposition in the marshy lands		- 02
ii)	c – cooling equipment/ by using aerosole so Persist in the environment for a long time. Accumulate in the body of organisms along	period	- 03
(C) i.	Widely dispersed in the environment Highly toxic reduce pests/ minimizing the formation of	` ,	- 02
	minimizing the spreading of diseases To successfully control the coconut caterpi		- 01
iii Use	(Promechotheca cumingi) / using BTI le e of orgarnic fertilizer / Reforestation for env Management / carbon foot print and shorte	vironmental balance/ waste	- 01
	methods and technology	_	- 01
	methods and teemfology		15
02 (A) i c	carbon, hydrogen, oxygen, nitrogen (For 03 c	or 04) - 01	15
	Enzymes	-	- 01
	Amylase/ Ptayalin		- 01
	(a) Cell wall/chloroplasts/ A large vacuole		- 01
	(b) Cell wall – to maintain the shape of the of the cell Larg vacuole – support / water balance	e cell/ support and protection	
	Chloroplasts- photosynthesis (For one	e fact)	- 01
V.	(a) KOH / potassium hydroxide		- 01
	(b) Absorbing carbon di oxide		- 01
	(c) rising of the coloured water through the	ie glass tube	- 01
. ,	Muscle tissue		- 01
	Skeletal muscles/ Heart muscles/ smooth		- 01
	Epithelial tissue		- 01
iv.			
	Moisturizing/ Humidifying inhaled air	r	
	• Warming up of inhaled air up to body t	emperature	
	• Removal of foreign matter from inhale	d air (For 2 facts)	- 02

		v.	(a) Ultra fitration/ selective reabsorption/ secretion -	01
			(b) Urea/ Uric acid/ kreatin	01
				15
03.	(A)	i.	(a) A	01
				01
		ii.		01
				01
		iv.		01
			(b) -	02
			H H	
		v.	(Electron distribution is not symmetrical in the bond of the molecule so	
			polarization occurs)	
			Because of polarization intermolecular bonds are formed. Due to intermolecular bonds high boiling point and high specific heat capacity occur	03
	(-)		1 / 1 / 1	02
	(B)		,	01
		ii.		01
			(b) When the temperature is high/low solubility of sugar is high/low	01
4.	(A)	;	The acceleration of a body is directly proportional to the unbalanced force	15
т.	(A)	1.		01
		ii.	J 1 1	01
	(B)	iv.	1 200 N F = ma 1 560 N - 1 500 N = 400 kg x a - $a = 60/400$ m s ⁻² = 0.15 m s ⁻² (without unit -00)	· 01 · 01 · 01 · 01

ii. Acceleration = Change of velocity

Time $0.15 = (v - 0) / 20 \text{ s} \quad \text{for applying} \qquad -01$ $v = 3 \text{ m s}^{-1} \quad \text{(with unit)} \qquad -01$ iii. Difference in displacement = $20 \text{ s} \times 3 \text{ m s}^{-1} = 60 \text{ m}$

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(C) i. Quantity of heat = 200 \times 4200 \text{ J kg}^{-1} \circ \text{C}^{-1} \times (100 \circ \text{C} - 30 \circ \text{C})
                                                                                                   - 01
                                 1000
                          = 58800 I
                                                                                                   - 01
        ii. Water changes its state from water (liquid) to water vapour at that point
                                                                                                   - 01
        iii. Power of the thermal coil = 58800 \text{ J} / 2 \times 60 \text{ s}
                                                                                                   - 01
                                      = 490 W
                                                                                                   - 01
                                                                                                   15
5. (A) i. Plantae
                                                                                                   - 01
        ii. Dicotyledonae (Dicots)
                                                                                                   - 01
        iii. Cycas/Pinus
                                                                                                   - 01
        iv. Amphibians
                                                                                                   - 01
    (B) i. anther, filament
                                                                                                   - 02
                                                                                                   - 02
        ii. (a) Ovary (b) Uterus
        iii. Progesterone
                                                                                                   - 01
        iv. (a) The combination of a gene pair for a particular character
                                                                                                   - 02
            (b) The genes that present in the same chromosome which are not segregated
                Independently
                                                                                                   - 02
    C) i. (a) absorbing water
                                                                                                   - 01
            (b) production of bile/converting excess glucose to glycogen
                                                                                                   - 01
        ii. Ptayalin/salivary amilase
        iii. Diastole/ Atrial contraction
                                                                                                   - 01
            Systole/ Ventricular contraction
                                                                                                   -01
        iv. X - Sensory Neuron
                                                                                                   - 01
            y – motor Neuron Brain(in cranial reflexes)
                                                                                                   <u>- 01</u>
                                                                                                   20
                                                                                                   - 01
06 (A) i.
           Colour change/ Blue colour solution becomes colourless
        ii.

    Take equal volume of copper sulphate with equal concentration to two equal

            test tubes
          • Take equal amount of Mg in the same physical condition and put in to the two
            test tubes seperately
          • Keep one test tube in a beaker with 0°C water and the other test tube in
            a beaker with 100°C water at the same time and observe
                                                                                                   - 03
                CuSO_4(aq) + Mg(s) \longrightarrow MgSO_4(aq) + Cu(s)
                                                                                                   - 03
          iii.
            (for the correct equation - 02 for the correct physical states - 01)
                                                                                                   - 01
        iv. Single displacement
    (B) i. Mass of an atomic mole of Magnesum= 24g
                                                                                                   - 02
        ii. Mass of an atom of Mg = 24 / 6 \times 10^{23} = 4 \times 10^{-23} g
                                                                                                   - 02
iii. Number of atomic moles= Mass/ Molor Mass = 6g
                                                                                                   - 01
                                                                                                   - 02
                                         0.25 mol
        iv. C
               = m
                   vΜ
                                                                                                   - 01
            m = CvM
                = 0.1 \text{ mol dm}^{-3} \times 100 \text{ dm}^{-3} \times 249.5 \text{ mol}^{-1}
                                                                                                   - 01
                           1000
                = 2.495 g
                = 2.5 g - (01 mark)
                                                                                                   - 03
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20

7. (A) i.	100 N + 100N = 200N	- 0
	ii.	To decrease the amount of force applied by a single person	
		To change the direction of force	- 0
	iii.	(a). In Equilibrium / balanced	- 0
		(b).	
	•	The three forces must be coplanar	
		The resultant of any two of the forces should be equal to the third force	
		in magnitude and opposite in direction	- 0
	iv	(a). $\frac{50}{100} \times 200 \text{ N} = 100 \text{ Nm}$	- 0
	1 V .	(a). 100 × 200 W = 100 WH	U
		(b). Clockwise	- 0
	v.	(a). $50 \text{ N} - (15 + 15 \text{ N}) = 20 \text{ N}$	- 0
		(b). $P = \pi + h dg / \pi + h p g$	- 0
			- 0
		(c). $P = 4 \text{ m } \times 1000 \text{ kg m}^{-3} \times 10 \text{ ms}^{-1} = 40000 \text{ N m}^{-2}$	- 0
ſB) i.	(a) centre of curvature	
(2	<i>,</i>	(b) focul point	
		(c) pole	- 0
	ii.	upright/ virtual/ large(Magnification more than 01)	- 0
		By dental doctors to examing teeth of patients/ to see the face when shaving	·
		the beard	- C
			20
8. (A) i	Natural classification	- 0
, o. (11	ii.	(a) Archaea	- 0
		(b) polar ice caps/volcanoes/ocean beds/deep inside the earth/	
		hot water spring	- 0
		Amphibians, Pisces, Arthropoda, Mollusca	- 0
	iv.	(a) x – parameceum y – chlamidomonas z – Amoeba	- 0 - 0
(B)	i.	(b) Hydrilla/ Keketiya/ valisnaria(a) Ultra sound waves	- C
(D)	1.	(b) To mix xhoxolate/ to scan internal organs of the human body/ to detect	Ü
		fine fractures in boilers, air planes/ to find the depth of the sea/	
		to solder metals	- 0
	ii.	Real depth - 14 N Apparent depth - 4 N	- 0
	111.	Density of lead is more relative to the density of water. Density of Styrofoam	
		is less relative to the density of water. Therefore lead balls sink in water Styrofoam balls float in water	- 0
	iv.	(a) Increases	- 0
		(b) Opticals fibres/ cutting of gems/ optical telephone wires/ In binoculars	<u>- 0</u>
			20
09. (A) i.	Fractional distilleration	- 0
٠٠. (ii.	Exothermic	- 0
	iii.		
		2C H + 2E O	-0
		$\frac{2C_8H_{18} + 25 O_2}{I}$	
		16CO ₂ + 18H ₂ O	

- 02 (B) i. Ethene ii. Ethane – no double bonds, Ethene – has double bonds - 02 - 02 iii. $F \subset C = C \subset F$ - 02 (C) i. Rectifying diodes - 02 ii. d, b iii. Indicator deflects in the same direction - 02 - <u>03</u> iv. 20